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- Q.1 Which one of the following parameters is very important for the non-ideal behavior of gases?**
- a. low pressure and low volume
 - b. high volume and high temperature
 - c. low temperature and high pressure
 - d. high temperature, high volume and high pressure
- Q.2 When a graph is plotted between pressure on x-axis and the PV/RT on y-axis for an ideal gas, then:**
- a. Hyperbolic curve is obtained
 - b. A straight line is obtained running in-between x-axis and y-axis
 - c. A straight line is obtained running parallel to pressure axis
 - d. A peak is obtained running parallel to compressibility factor axis
- Q.3 Which one of the following gases show more deviation from ideal behavior?**
- a. CO_2
 - b. H_2
 - c. N_2
 - d. He
- Q.4 At what temp the gases is likely to obey the K.M.T?**
- a. low
 - b. highest
 - c. high
 - d. lowest
- Q.5 Charle's law is an example of _____ process:**
- a. isothermal
 - b. isobaric
 - c. isochoric
 - d. All of given
- Q.6 Which of the following is an example of polar molecular solid**
- A. Naphthalene
 - B. Dry ice
 - C. Iodine
 - D. Sugar

- Q.7 Solid CO₂ (Dry ice) is an example of _____ crystals**
- A. Covalent
 - B. Ionic
 - C. Metallic
 - D. Molecular
- Q.8 If pressure remains constant at given temp the volume of an ideal gas is doubled as compared to volume at 0°C:**
- a. -273°C
 - b. 273K
 - c. 273°C
 - d. 546°C
- Q.9 Correct order of boiling points of the given liquids is:**
- A) H₂O > HF > HCl > NH₃
 - B) HF > H₂O > HCl > NH₃
 - C) H₂O > HF > NH₃ > HCl
 - D) HF > H₂O > NH₃ > HCl
- Q.10 Which of the following sets of solid elements A, B, C and D includes a giant metallic structure, a macromolecular structure and a simple molecular structure?**
- A) Na, Mg, Al, S
 - B) C, Si, Sn, Pb
 - C) Al, Si, S, P
 - D) Al, S, Si, CO₂
- Q.11 Which is correct increasing order of boiling points of the hydrides of VI-A group?**
- a. H₂Te < H₂S < H₂Se < H₂O
 - b. H₂S < H₂Se < H₂Te < H₂O
 - c. H₂Se < H₂S < H₂O < H₂Te
 - d. H₂Te < H₂Se < H₂S < H₂
- Q.12 The overall structure of ice is just like that of a:**
- a. Graphite
 - b. Sugar
 - c. Tin
 - d. Diamond

Q.13 Which one of the following statement is not true about evaporation?

- a. Spontaneous process
- b. Causes cooling
- c. Temperature increasing
- d. Increases with increase in surface area

Q.14 Which one of the following halogen has highest melting and boiling point?

- a. Br_2
- b. I_2
- c. Cl_2
- d. F_2

Q.15 Which statement about molecular solids is correct?

- a. They contain a network of atoms
- b. They have high melting points
- c. They are very hard and greater energy is required to break them
- d. They volatility is very high

Q.16 Covalent crystals are bad conductor of electricity due to absence of free electrons and ions except:

- a. Silicon Carbide
- b. Graphite
- c. Cadmium iodide
- d. Born nitride

Q.17 The covalent crystals having giant molecules like diamond and silicon carbide are:

- a. Soluble in all the solvents
- b. Insoluble in all the polar solvents only
- c. Soluble in all the non-polar solvents only
- d. Insoluble in all the solvents

Q.18 Which of the following statements about ionic solids, covalent solids and molecular solids is incorrect:

Options	Properties	Ionic solids	Covalent solids	Molecular solids
A)	Example	NaCl. CaO	Diamond. SiC	I ₂ , CO ₂ , HCl, Ice
B)	Basic component	Ions	Atom	Molecule
C)	Electrical conductivity	Non-conductor in solid state	Non-conductor except graphite	Non-conductor except HCl in H ₂ O
D)	M.P and B.P	Very high M.Ps and B.Ps	Very low M.Ps and B.Ps	Very high M.Ps and B.Ps

Q.19 Identify the amorphous solid :

- a. Sugar
- b. Cellulose
- c. Iodine
- d. none of given

Q.20 Which one of the following pairs contains polar molecular solids?

- a. Iodine and Sugar
- b. Carbon dioxide and Ice
- c. Phosphorus and Carbon dioxide
- d. Sugar and Ice

Q.21 Identify incorrect statement:

Options	Types	Properties
A)	Metallic solids	Malleable and ductile
B)	Ionic solids	Hard and brittle
C)	Covalent solids	Higher melting and boiling points
D)	Molecular solids	Their basic unit is atom

Q.22 When water freezes at 0°C, its density decrease due to:

- A) Cubic structure of ice
- B) Empty spaces present in the structure of ice
- C) Change of bond lengths
- D) Change of bond angles

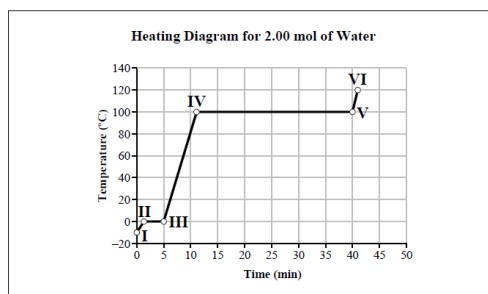
Q.23 Diamond, silicon carbide and silicon (IV) oxide have all of the following properties EXCEPT:

- A) Called network solids
- B) They are with open structure
- C) Melting point of silicon carbide (SiC) is higher than diamond
- D) Powdered diamond and silicon carbide are used as abrasive

Q.24 Which of the following statements about ionic solids, covalent solids and molecular solids is incorrect:

Options	Properties	Ionic solids	Covalent solids	Molecular solids
A)	M.P and B.P	High	Very high	Low
B)	Electrical conductivity	Non-conductors when solid	Non-conductors EXCEPT graphite	Non-conductor EXCEPT HCl in water
C)	Basic bonding	Ionic bond	Covalent bond	Hydrogen bonding, dipole dipole forces. London dispersion forces
D)	Solubility	Soluble in polar solvents	Rapidly soluble in polar and non-polar solvent	- Polar compounds in polar solvents - Non polar compounds in non-polar solvents

Q.25 The region on the diagram where liquid water is undergoing a change primarily in kinetic energy is between



- a. II and III
- b. V and VI
- c. IV and V
- d. III and IV

Q.26 Correct gas equation is

- A) $\frac{V_1 T_2}{P_1} = \frac{V_2 T_1}{P_2}$
- B) $\frac{P_1 T_1}{V_1} = \frac{P_2 V_2}{T_2}$
- C) $\frac{P_1 V_1}{P_2 V_2} = \frac{T_1}{T_2}$
- D) $\frac{V_1 V_2}{T_1 T_2} = P_1 P_2$

Q.27 By what factor does the average velocity of a gaseous molecule increase when the temperature (in Kelvin) is doubled?

- A) 2.0
- B) 4.0
- C) 2.8
- D) 1.4

Q.28 The basic distinction between solids, liquids and gases lies in difference between:

- A. Strength of the bonds
- B. Size of molecules
- C. Space which the molecules occupy
- D. All of above

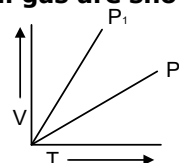
Q.29 Under what conditions the density of an ideal gas will be maximum at

- A. High temperature and low pressure
- B. Low temperature and high pressure
- C. Low temperature and low pressure
- D. When temperature and pressure are same

Q.30 The correct value of general gas constant R is

- A.** $8.314 \text{ atm dm}^3 \text{ mol}^{-1} \text{ K}^{-1}$
- B.** $0.0821 \text{ Jmol}^{-1} \text{ K}^{-1}$
- C.** $62.4 \text{ torr cm}^3 \text{ mol}^{-1} \text{ K}^{-1}$
- D.** $8.314 \times 10^7 \text{ erg mol}^{-1} \text{ K}^{-1}$

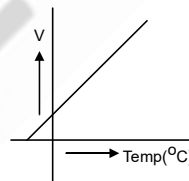
Q.31 V vs T curves at constant pressure P₁ and P₂ for an ideal gas are shown below.



Which of the following is correct?

- A)** $P_1 > P_2$
- B)** $P_1 < P_2$
- C)** $P_1 = P_2$
- D)** All of above

Q.32 The following graph illustrates



- A)** Boyle's law
- B)** Charles law
- C)** Dalton's law
- D)** Gay-Lussac's law

Q.33 Which of the following has strongest intermolecular forces of attraction

- A.** Hydrogen(H_2)
- B.** Chlorine (Cl_2)
- C.** Iodine (I_2)
- D.** Methane (CH_4)

Q.34 The forces which are present between the ions and water molecules are known as

- A.** Dipole induced dipole forces
- B.** Dipole-dipole forces
- C.** Ion dipole forces
- D.** London dispersion forces

- Q.35 In which compound hydrogen bonding is not present?**
A. H_2O
B. CH_4
C. NH_3
D. C_2H_5OH
- Q.36 Which has strongest bonding in the solid state?**
A. Hydrogen Chloride (HCl)
B. Xenon(Xe)
C. Chlorine (Cl_2)
D. Sodium Chloride (NaCl)
- Q.37 The spontaneous mixing of particles is called:**
A. **evaporation**
B. **sublimation**
C. **diffusion**
D. **boiling**
- Q.38 Liquids have no definite shape; it is because**
A. The molecules of liquid are in constant motion by sliding over each other
B. The intermolecular forces of liquids are weaker than gases
C. The liquid molecules have kinetic energy less than solids
D. All statements are correct
- Q.39 Propanone is miscible in water due to**
A. Both are polar molecules
B. Dipole-dipole attraction between them
C. Hydrogen bonding between them
D. All of these
- Q.40 The strongest H-bond is**
A. F---H
B. O----H
C. N----H
D. O----F